

## Pure

A substance that consists of only one element or only one compound.



## Impure

A substance that consists of more than one element or compound.



## Mixture

Two or more substances that are not joined together. The substances can be elements, compounds, or both.

## Melting point

The temperature at which a solid changes into a liquid as it is heated.



## Formulation

A mixture that is always made with the same proportions of the same substances.



## Chromatogram

The results of separating mixtures by chromatography.

## Soluble

Able to dissolve in solvent. For example, sugar is soluble in water because it dissolves to form sugar solution.



## Mobile phase

Phase in chromatography that moves, usually a solvent or mixture of solvents.



## Solvent

The liquid in which the solute dissolves to form a solution.

## Stationary phase

Phase in chromatography that does not move, for instance, the paper in chromatography.



## Dissolved

A substance is said to be dissolved when it breaks up and mixes completely with a solvent to produce a solution.



## Ion

Electrically charged particle, formed when an atom or molecule gains or loses electrons.

## Precipitate

A suspension of particles in a liquid formed when a dissolved substance reacts to form an insoluble substance, eg in a precipitation reaction.



## Limewater

Calcium hydroxide solution. It turns milky in the presence of carbon dioxide.



## Halide

A halide ion is an ion formed when a halogen atom (an atom from group 7) gains one electron. Halide ions have a single negative charge. Ionic compounds containing halide ions may be called halides.