

Reactivity

A measure of how vigorously a substance will react. The more reactive it is, the greater its reactivity and the more vigorous its reactions will be.



Oxidation

The gain of oxygen, or loss of electrons, by a substance during a chemical reaction.



Reduction

The loss of oxygen, gain of electrons, or gain of hydrogen by a substance during a chemical reaction.

Displacement reaction

A reaction that occurs when a more reactive element replaces a less reactive element in a compound.



pH

Scale of acidity or alkalinity. A pH (power of hydrogen) value below 7 is acidic, a pH value above 7 is alkaline.



Neutralisation

The reaction between an acid and a base to form a salt plus water.

Base

A substance that reacts with an acid to neutralise it and produce a salt.



Alkali

A base which is soluble in water.



Titration

A quantitative procedure in which two solutions react in a known ratio, so if the concentration of one solution is known and the volumes of both are measured, the concentration of the other solution can be determined.

End-point

In a titration, the point at which the indicator first permanently changes colour.



Concentration

A measure of the mass or amount of solute dissolved in a given volume of solvent or solution.



Electrolysis

Decomposition (of a liquid electrolyte) using (a direct current of) electricity.

Fold

Electrolyte

Liquid which conducts electricity and is decomposed by it.



Anode

Positive electrode.



Cathode

Negative electrode.

B B C BITESIZE

First cut along the solid lines and then fold each flashcard along the dotted line, so the keyword is on the front and the explanation is on the back.